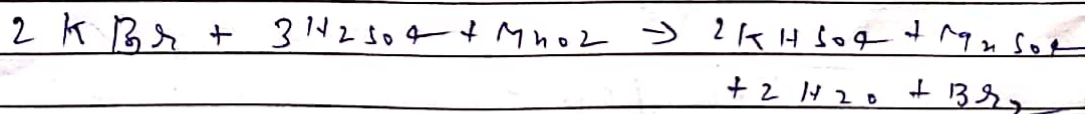
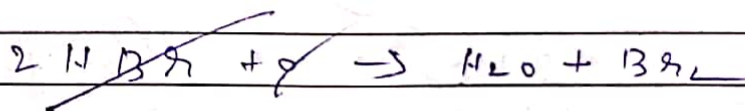
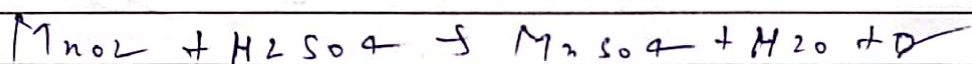
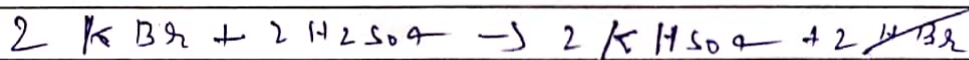


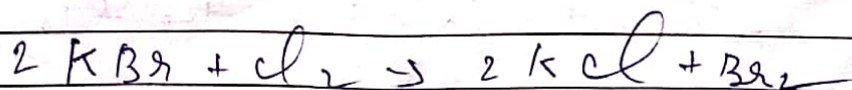
BROMINE: Bromine (Br_2) exist in liquid form with yellow colour. It causes injury on skin, therefore use of liquid Br_2 in laboratory is done with Caution.

Br_2 is prepared as follows.

Reaction of KBr , MnO_2 and Conc H_2SO_4 on heating gives Br_2

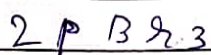
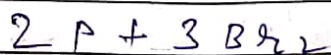


(ii) Reaction of KBr with Cl_2 gives Br_2



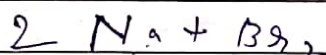
Properties of Br_2

(a) Non-Metals such as Phosphorus reacts with Bromine and forms Phosphorus tribromide.

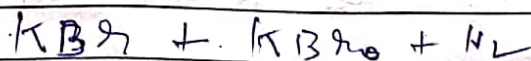
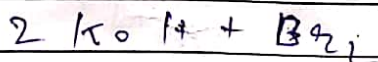


Phosphorus tribromide

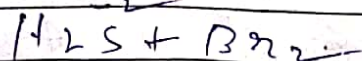
(b) Bromine reacts with Metal such as Sodium and forms Sodium Bromide



(c) Br_2 reacts with Dilute KOH and forms Bromide and Hypobromite



(d) Bromine oxidises H_2S (Hydrogen sulphide) into Sulphur



In this reaction

oxidation no. of

'S' changes -2 to '0'

means oxidation takes place.