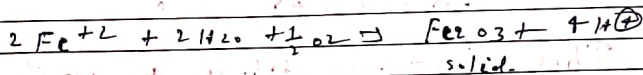
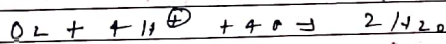
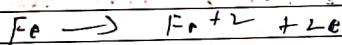


TOPIC: CORROSION.

Corrosion is a process in which metal is oxidised by losing electron.

lost electron is accepted by oxygen and oxide ions are formed.

Iron is rusted in presence of moisture and oxygen.



Corrosion can be avoided by adopting certain process.

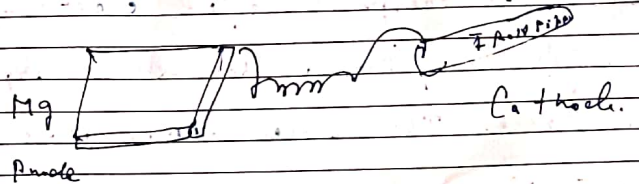
(A) Galvanisation:- In this process Zinc is coated over iron to protect it from corrosion.

CATHODIC PROTECTION:-

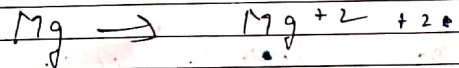
In this process more reactive metal is made as sacrificial

ANODE.

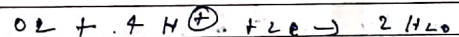
More reactive metal such as zinc is buried under earth.



IRON PIPE is connected to Mg by wire.



Mg acts as Anode and Iron acts as Cathode. Minerals in soil acts as electrolytes.



The reactive metal is sacrificed to protect the iron. Reactive metal requires time to time replace most